

	principle quantum number	Azimuthal quantum number	Magnetic quantum number	Spin quantum numbers	set of quantum numbers	orbital
	$n =$	$l =$	m_l	$+1/2$ or $-1/2$		
		0 to $n-1$	negative l to positive l			
1st electron	1	0	0	$+1/2$	1,0,0,+1/2	1s
2nd electron	1	0	0	$-1/2$	1.0.0.-1/2	1s
3rd electron	2	0	0	$+1/2$	2.0.0.+1/2	2s
4th electron	2	0	0	$-1/2$	2.0.0.-1/2	2s

5th electron	2	1	-1	+1/2	2,1,-1,+1/2	2px
6th electron	2	1	-1	-1/2	2,1,-1,-1/2	2px
7th electron	2	1	0	+1/2	2,1,0,+1/2	2py
8th electron	2	1	0	-1/2	2,1,0,-1/2	2py
9th electron	2	1	+1	+1/2	2,1,1,+1/2	2pz
10th electron	2	1	+1	-1/2	2,1,1,-1/2	2pz